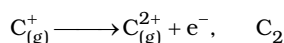
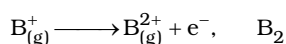
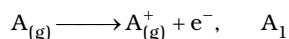


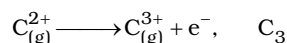
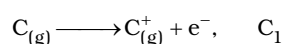
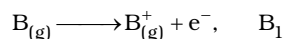
Date Planned : __ / __ / __	Daily Tutorial Sheet - 9	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level - 2	Exact Duration : _____

106. Consider the following ionisation reactions:

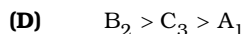
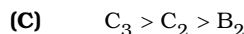
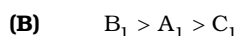
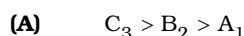
I.E. (kJ mol^{-1})



I.E. (kJ mol^{-1})



If monovalent positive ion of A, divalent positive ion of B and trivalent positive ion of C have zero electron. Then incorrect order of corresponding I.E. is:



107. The incorrect statement is:

(A) The second ionisation energy of Se is greater than that of second ionisation energy of As

(B) The first ionisation energy of C^{2+} ion is greater than the first ionization energy of N^{2+} ion

(C) The third ionisation energy of F is greater than the third ionisation energy of O

(D) Halogens have highest I.E. in respective period

108. First three ionisation energies (in kJ/mol) of three representative elements given below:

Element	IE_1	IE_2	IE_3
P	495.8	4562	6910
Q	737.7	1451	7733
R	577.5	1817	2745

Then incorrect option is:

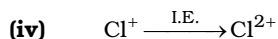
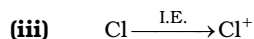
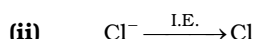
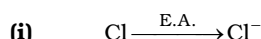
(A) Q : Alkaline earth metal

(B) P: Alkali metals

(C) R: s-block element

(D) They belong to same period

109. Which of the following statement is correct regarding following process?



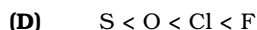
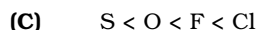
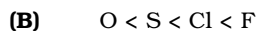
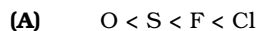
(A) $|\text{I.E. of process (ii)}| = |\text{E.A. of process (i)}|$

(B) $|\text{I.E. of process (iii)}| = |\text{I.E. of process (ii)}|$

(C) $|\text{I.E. of process (iv)}| = |\text{E.A. of process (i)}|$

(D) $|\text{I.E. of process (iv)}| = |\text{I.E. of process (iii)}|$

110. The correct order of increasing electron affinity of the following elements is:



111. The second electron gain enthalpies (in kJ mol^{-1}) of oxygen and sulphur respectively are:

